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Forbes ,1979

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equation pressure
remains equal to the
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the piston moves up as
long as both phases
remain present all that
happens is that more
water turns to steam and
the heat reservoir must
supply the latent heat
of vaporization λ 40 65
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11 5 vapor pressure
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curve of water
intersects the p 1000
mmhg line at about 110 c
this is therefore the
boiling point of water
at 1000 mmhg b the
vertical line
corresponding to 250 c
intersects the vapor
pressure curve of
mercury at p 75 mmhg
hence this is the
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liquid is defined as the
pressure exerted by the
molecules that escape
from the liquid to form
a separate vapor phase
above the liquid surface
this pressure is formed
in a thermodynamic
equilibrium state in a
closed container at a
certain temperature
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phase thermodynamics to
be curve fitted to 300k
about 142k lower than the
normal boiling point 542k
2 of kerosene

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