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calorimeter th water in it and the hot object made up of copper which is
put in the calorimeter are the same the initial temperature of the
calorimeter and water is 30 c and that of hot object is 60 c the specific
heat of copper and water are 0 09 cal gm c and 1 cal gm c respectively
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web oct 28 2019 solution use this equation $q = \text{specific heat} \times m \times \Delta t$
where q is heat flow m is mass in grams and Δt is the temperature
change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ j}$
 $\text{g} \times 110 \text{ g} \times 26.6 \text{ c} - 25.0 \text{ c}$ $q_{\text{water}} = 550 \text{ j}$ $q_{\text{water}} = 550 \text{ j}$
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 what is the specific heat of aluminum if the temperature of a 28 4 g
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 and calculations because the heat released or absorbed at constant
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 $\Delta H_{rxn} = q_{rxn} / m_{cs} \Delta t$ 7 3 14 7 3 14 $\Delta H_{rxn} = q_{rxn} / m_{cs} \Delta t$
 m a t e r m c s Δt

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 i have 125 grams of water at a temperature of 25 0 c and heat it to 145 0 c
 what will the change in enthalpy be 2 sketch the heating curve
 temperature vs time for the process in 1 3 i have 18 7 grams of an
 unknown liquid chemical at a temperature of 35 0 c

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 25 0 ml of 1 00 m hcl at 25 0 c is added to 25 0 ml of 1 00 m naoh at 25
 0 c in a foam cup calorimeter a reaction occurs and the temperature rises
 to 32 0 c calculate the enthalpy change in kj for this reaction assume the

densities of the solutions are 1 00 g ml and that

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 quantity of heat transferred

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 with an excess of oxygen and ch₄ in a bomb calorimeter with a total heat

capacity of 15.48 kJ °C the temperature of the calorimeter and its contents changes from 19.77 °C to 33.90 °C calculate q and h for the following reaction as written 2 NH₃(g) + 3 O₂(g) → 2 H₂O(l)

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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web 6 a 218 g sample of steam at 121 °C is cooled to ice at 14 °C find the change in heat content of the system 7 in going from ice at 34 °C to steam at 138 °C a sample of water absorbs 1.41 × 10⁵ J find the mass of the sample answers 1 2.68 × 10⁴ J 2 6.23 × 10⁴ J 3 1.71 × 10⁴ J 4 1.11 × 10⁵ J 5 2.11 × 10⁵ J 6

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