

Avrpal Net 2100

Ensheng Dong

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this laboratory

exercise relies on a

titration technique

to determine an

unknown

concentration of

monoprotic acid in

solution in the

process of titration

a basic solution is

gradually added to

the acidic solution

until complete

neutralization is

obtained

acids bases salts

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acid and strong

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strong acids and

strong bases refer

to species that

completely

dissociate to form

ions in solution by

contrast weak acids

and bases ionize

only partially and

the ionization

reaction is

reversible

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used in an acid

base experiment in

order to determine

the concentrations

of solutions of acids

and bases through

the titration

process we are able

to identify physical

changes to the

mixture such as the

colour change to

indicate the end

point of the

experiment

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are usually used to

find the amount of a

known acidic or

basic substance

through acid base

reactions the

analyte titrand is

the solution with an

unknown molarity

the reagent titrant

is the solution with

a known molarity

that will react with

the analyte

acid base

titrations

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web conclusion the

purpose of the

experiment was to

identify the

unknown solid acid

17 by calculating its

molar mass this was

done by titration of

the acid with

sodium hydroxide

in order to find the

volume of sodium

hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration

data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to

determine the standardization of sodium hydroxide naoh and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution
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 web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown

base key terms
 titration a process where a solution of known strength is added to a certain volume of a treated sample containing an
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 web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4$ aq $2\text{NaOH aq Na}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction
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