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TD Snyder

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below 4 no g 6 h 2 o l 4 nh 3

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reaction endothermic or

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gt 0 b how much heat will be

absorbed or released if 10 g

of no is reacted with 10 g h

2 o

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the entire reaction was

multiplied by 2 therefore Δh

must also be halved 3

answer c explanation option

c is incorrect because 566 kj

mol is the Δh per mole of O_2 and not for 2 moles even though options b and d are expressed in grams they are still correct because 56
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 $2H_2(g) + O_2(g) \rightarrow 2H_2O(l)$ energy endothermic the transfer of heat from the surroundings to the system energy
 $2HgO(s) \rightarrow 2Hg(l) + O_2(g)$
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a a thermodynamic system has two means of exchanging energy with its surroundings heat flow and pressure volume work heat flow q is the transfer of thermal energy due to temperature difference work w is macroscopic energy transfer between the system and its surroundings that occurs

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