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web question 1 find the mole fraction of methanol  $CH_3OH$  and water in a solution prepared by dissolving 45 g of alcohol in 40 g of  $H_2O$  molar mass of  $H_2O$  is 18 gm mole and molar mass of  $CH_3OH$  is 32 gm mole solution moles of  $CH_3OH$   $\frac{45}{32} = 1.406$  moles of  $H_2O$   $\frac{40}{18} = 2.22$  moles therefore according to the equation

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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction times the total pressure  $p_a = x_a p_t$  mole fraction  $x_a = \frac{\text{moles of } a}{\text{total moles}}$   $x_a = \frac{n_a}{n_{\text{total}}}$  1 all  $x_n = 1$  mole fraction of substance  $a$  is often denoted by  $x_a$  and *mole fraction definition formula and solved examples byju s* - Sep 14 2023 web 1 in the given mixture the sum of all the mole fractions is equal to one  $x_a + x_b + \dots = 1$  2 the multiplication of the mole fraction by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction mole fraction is not dependent on the temperature

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web calculate the mass percent mole fraction molality and molarity of hexane  $C_6H_{14} = 660 \text{ g/mol}$  3 when it is mixed with heptane  $C_7H_{16} = 684 \text{ g/mol}$  3 in a 35.0 : 55.0 ml ratio assume that the volumes are additive mole fraction chemteam - Oct 15 2023

web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web jan 8 2022  $\sum x_i = 1$  the sum of all of the mole percents of a mixture equals 100 mole fraction properties and advantages the mole fraction offers advantages over some of the other units of concentration unlike molarity mole fraction is

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web since the mole fraction of the o<sub>2</sub> is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.21 0.050 4.2 atm

problem 9 a sample of 1.43 g of helium and an unweighed quantity of o<sub>2</sub> are mixed in a flask at room temperature

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