

Isotopes Decay Calculator

Kalyan K. Dutta Choudbury

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Calculation of Isotope Production by the Cyclic Irradiation of Thorium Robert G. Nisle,1959

Calculation of Equilibrium Constants for Isotopic Exchange Reactions Jacob Bigeleisen,Maria Goeppert Mayer,1947

Retention and Elimination of Radium Isotopes Produced by the Decay of Thorium Parents Within the Body John C. Reynolds,Philip F. Gustafson,Leonidas D. Marinelli,1957 Radium daughters are produced in vivo by radioactive decay of thorium isotopes in the body and partially eliminated by metabolic processes. If the assumption is made that the radium formed gains access to the general circulation and it is retained, it is possible to calculate the elimination of radium from the body as a function of time, and the radioactive equilibrium of radium to thorium in vivo. These calculations have been performed for the case of a single administration of thorium (masses 227, 228, 229, 230, and 230).

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College Physics Michael Tammaro,2018-12-18 Tammaro's College Physics, First Edition will convert more students from passive to active learners through a unique presentation of material built from the ground up in a digital environment. When students become active learners, they study smarter by spending time on content that will help them improve their understanding of key concepts (NOT skipping straight to the problems to find out what they don't know). College Physics, First Edition utilizes an assignable, module structure with frequent assessment check points at various difficulty levels to ensure maximum points of student engagement and retention.

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Radioactive Isotopes of Gadolinium R. E. Hein,Adolf Frank Voigt,1950

The Masses of Heavy Isotopes Martin O. Stern,1948

Alpha Decay in Isotopes of Atomic Number Less Than 83 Stanley Gerald Thompson,S. G. Thompson,1949 Some time ago we started work in an attempt to observe alpha-particle decay in isotopes of atomic number less than 83. In the first experiments, thin targets of gold leaf were bombarded with 190-Mev deuterons in the 184-inch cyclotron. Two alpha-decay periods were observed in these targets; one of 0.7 minutes half-life and another of 4.3 minutes half-life. The alpha-particle energies were 5.7 and 5.2 Mev, respectively. Chemical separations proved that the 4.3-minute period is due to a gold isotope and suggested that the 0.7-minute period is due to a mercury isotope. The mass numbers of these new isotopes have not been determined. However, the results of excitation-functions in the production of the gold isotope by bombarding gold and platinum with protons suggest that its mass number lies in the range 185-188. The work on this isotope indicates that the alpha to electron capture branching ratio is of the order of magnitude of 10-4, and that positron activity accompanies the 4.3-minute alpha-period.

Pharmaceutical Calculations Joel L. Zatz,Maria Glaucia Teixeira,2013-03-27 Pharmacists are required to make certain kinds of calculations that determine the quantities of materials required for filling prescriptions and making up formulas. The new and expanded topics introduced in the fourth edition teach pharmacists and pharmacy students how to do the calculations required in current practice, covering important areas such as handling injectibles, including those used in parenteral nutrition and radiopharmaceuticals. The book also includes new chapters on isotonicity, intravenous fluids, and nutritional calculations. Features: * New concepts introduced in sequence, encouraging the student to master each concept before moving ahead * Many examples and practice problems, all with answers and the availability of rapid feedback build confidence * Filled with practical instruction relevant to the problems pharmacist face in their practice

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