

Amanda Easton

Ian Pickup

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of the two possible iron sulfate compounds formed when reacting aqueous copper ii sulfate with solid iron metal the techniques used in this lab were weighing by difference quantitative transfer and vacuum filtration

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as

a byproduct of the process as iron ions the ion meta dissolves
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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation *lab report stoichiometry copper sulfate and iron 2023* - Dec 27 2021 web lab report stoichiometry copper sulfate and iron the oxidation of trimethylamine by copper ii chloride nov 10 2022 this research was on the irreversible oxidation of trimethylamine by copper ii chloride which takes place

at temperatures of about 75 and higher the coordination of trimethylamine by copper ii **lab report on iron stoichiometry 1098 words bartleby** - Jan 28 2022 web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron **stoichiometry lab copper sulfate iron pdf stoichiometry** - Jun 01

2022 web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4 **stoichiometry lab the reaction of iron with copper ii sulfate** - Oct 05 2022 web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much

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 how much reactant do you
 need to make that much
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 calculate the ratio of moles
 of reactant to moles of
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 will use stoichiometric
 principles to deduce the
 appropriate equation for the
 reaction between metallic
 iron and a solution of copper
 ii sulfate this reaction
 produces metallic copper
 which is seen precipitating
 as a finely divided red
 powder
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 web in this lab we will use
 experimental evidence and
 the principles of
 stoichiometry to determine
 which ion is formed in a
 single replacement reaction

between iron and copper ii
 sulfate pre lab assignment in
 preparation for the lab
 activity we will investigate
 another single replacement
 reaction between copper
 and silver nitrate
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 web through this experiment
 we sought to use
 stoichiometric principles
 which derive from
 stoichiometry the study of
 substance quantities in ratio
 or proportion in a chemical
 reaction in this lab
 experiment the principles
 help distinguish the
 equation between iron
 powder and aqueous copper

ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship

between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate *lab report 4 stoichiometry of iron copper ii sulfate* - Feb 09 2023

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iron ions dissolve in the copper ii sulfate solution to produce copper ii ions **stoichiometry reaction of iron with copper ii sulfate** - Apr 11 2023

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experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding

heated cuso₄ solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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